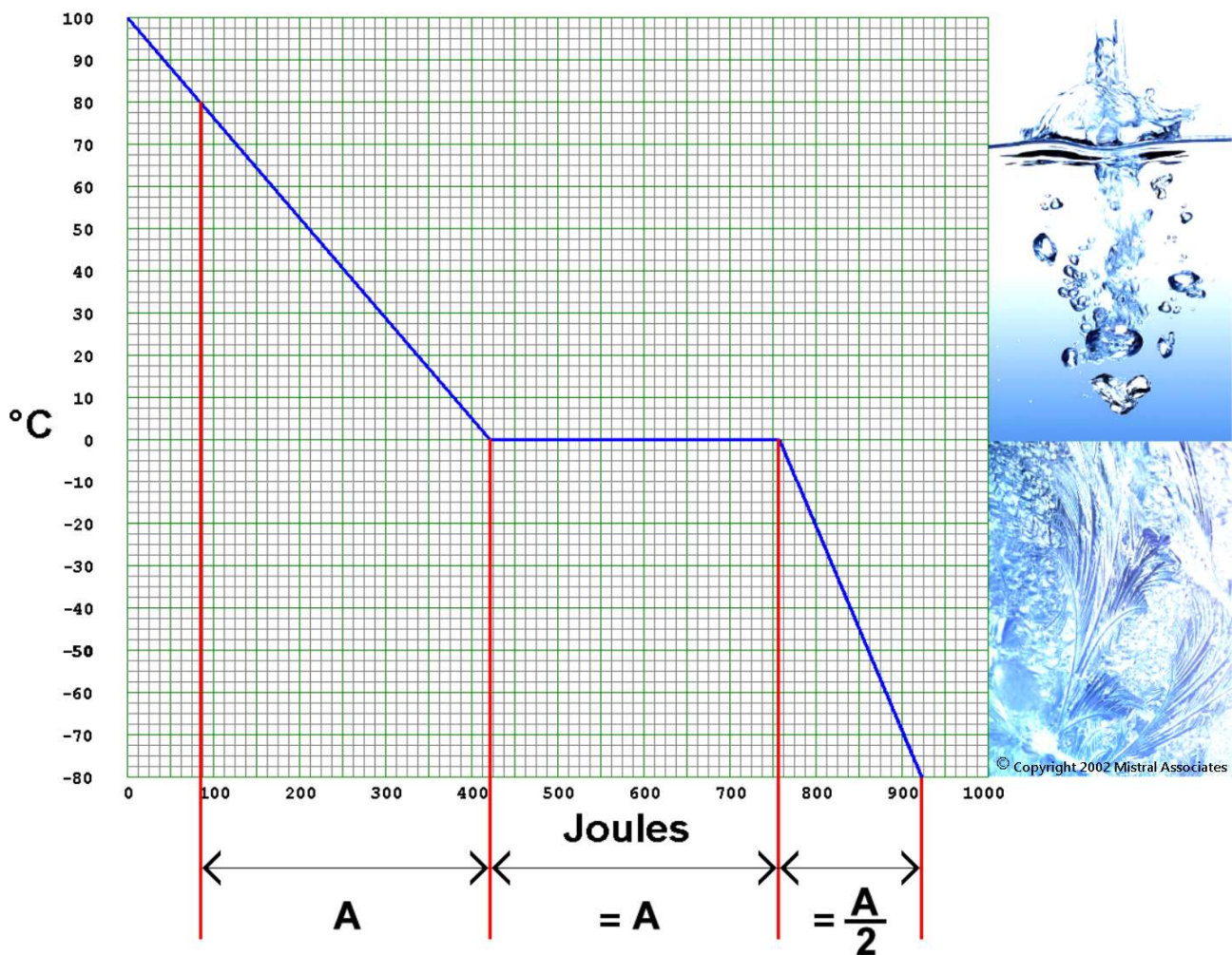


An example of how Latent Heat and Specific Heat works

It is necessary to remove around the same amount of energy from water at 0°C to remove its Latent Heat, and thus turn it into ice at 0°C, as it does to reduce its temperature from 80°C to 0°C.



To reduce the temperature of ice from 0°C to -80°C requires the removal of less than half the energy than it did to either remove the Latent Heat from the water or to reduce the temperature of the water by 80 degrees.

Mistral's commitment:

Bringing the benefits of computerisation to our industry - without the historically associated problems.

